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Chemistry
Standard level
Paper 2

Wednesday 18 May 2022 (afternoon)

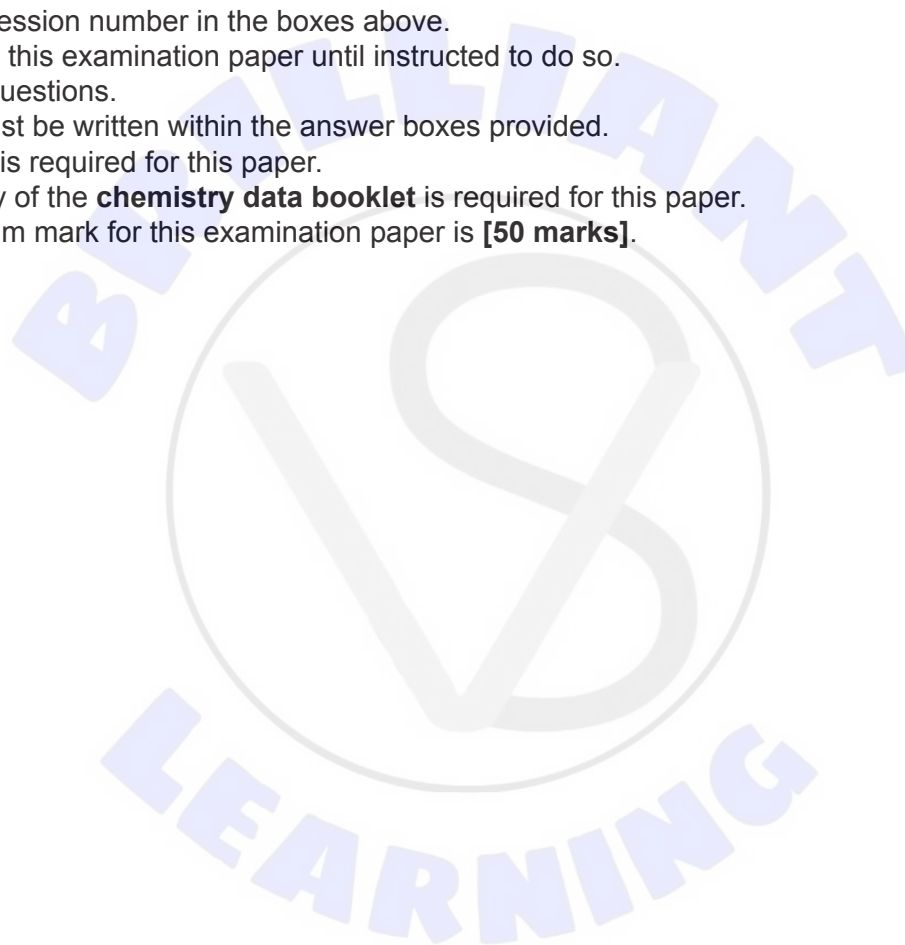
Candidate session number

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1 hour 15 minutes

Instructions to candidates

- Write your session number in the boxes above.
- Do not open this examination paper until instructed to do so.
- Answer all questions.
- Answers must be written within the answer boxes provided.
- A calculator is required for this paper.
- A clean copy of the **chemistry data booklet** is required for this paper.
- The maximum mark for this examination paper is **[50 marks]**.



Answer **all** questions. Answers must be written within the answer boxes provided.

1. When heated in air, magnesium ribbon reacts with oxygen to form magnesium oxide.

(a) (i) Write a balanced equation for the reaction that occurs. [1]

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(ii) State the block of the periodic table in which magnesium is located. [1]

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(iii) Identify a metal, in the same period as magnesium, that does **not** form a basic oxide. [1]

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(b) The reaction in (a)(i) was carried out in a crucible with a lid and the following data was recorded:

Mass of crucible and lid = 47.372 ± 0.001 g

Mass of crucible, lid and magnesium ribbon before heating = 53.726 ± 0.001 g

Mass of crucible, lid and product after heating = 56.941 ± 0.001 g

(i) Calculate the amount of magnesium, in mol, that was used. [1]

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(Question 1 continued)

- (ii) Determine the percentage uncertainty of the mass of product after heating. [2]

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- (iii) Assume the reaction in (a)(i) is the only one occurring and it goes to completion, but some product has been lost from the crucible. Deduce the percentage yield of magnesium oxide in the crucible. [2]

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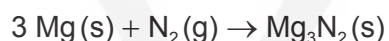
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- (c) When magnesium is burnt in air, some of it reacts with nitrogen to form magnesium nitride according to the equation:



- (i) Evaluate whether this, rather than the loss of product, could explain the yield found in (b)(iii). [1]

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- (ii) Suggest an explanation, other than product being lost from the crucible or reacting with nitrogen, that could explain the yield found in (b)(iii). [1]

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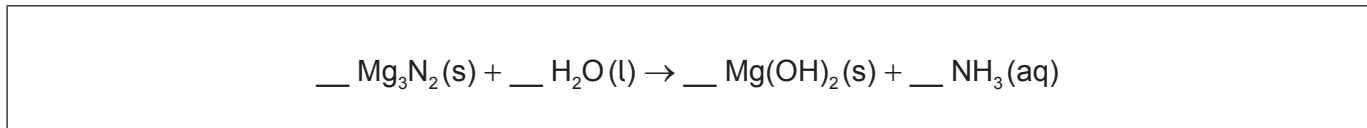
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(Question 1 continued)

(d) The presence of magnesium nitride can be demonstrated by adding water to the product. It is hydrolysed to form magnesium hydroxide and ammonia.

(i) Calculate coefficients that balance the equation for the following reaction. [1]



(ii) Determine the oxidation state of nitrogen in Mg_3N_2 and in NH_3 . [1]

Mg_3N_2 :

NH_3 :

(iii) Deduce, giving reasons, whether the reaction of magnesium nitride with water is an acid–base reaction, a redox reaction, neither or both. [2]

Acid–base: Yes No

Reason:

.....

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Redox: Yes No

Reason:

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(This question continues on the following page)



(Question 1 continued)

(e) Most nitride ions are $^{14}\text{N}^{3-}$.

(i) State the number of subatomic particles in this ion. [1]

Protons:
Neutrons:
Electrons:

(ii) Some nitride ions are $^{15}\text{N}^{3-}$. State the term that describes the relationship between $^{14}\text{N}^{3-}$ and $^{15}\text{N}^{3-}$. [1]

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(iii) The nitride ion and the magnesium ion are isoelectronic (they have the same electron configuration). Determine, giving a reason, which has the greater ionic radius. [1]

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(f) Suggest **two** reasons why atoms are no longer regarded as the indivisible units of matter. [2]

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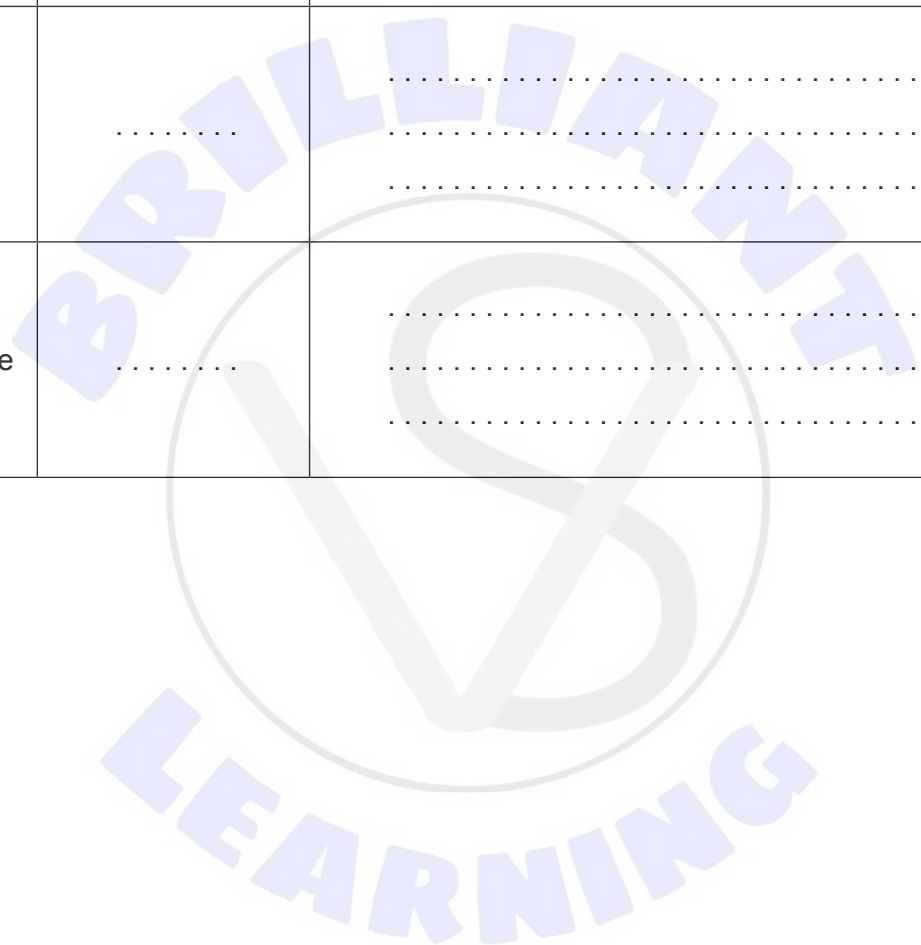


(Question 1 continued)

(g) State the types of bonding in magnesium, oxygen and magnesium oxide, and how the valence electrons produce these types of bonding.

[4]

Substance	Bond type	How the valence electrons produce these bonds
Magnesium
Oxygen
Magnesium oxide

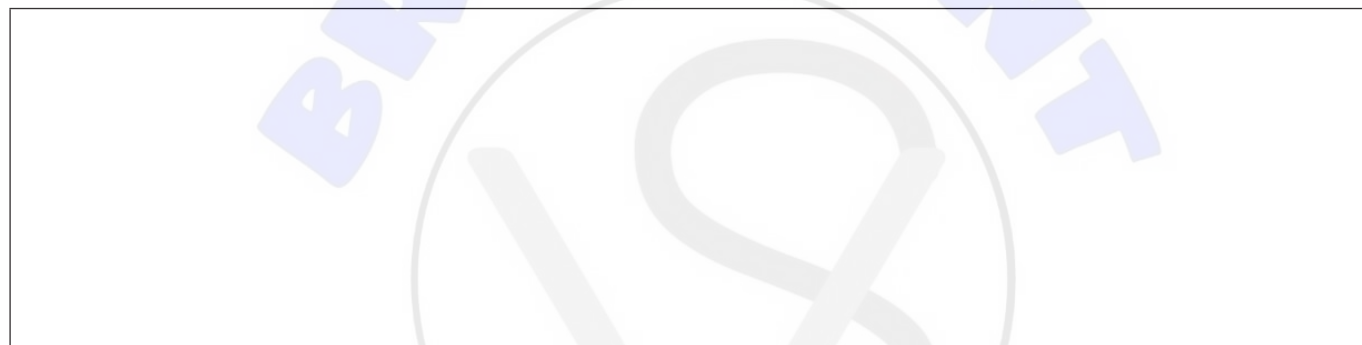


2. Ammonia, NH_3 , is industrially important for the manufacture of fertilizers, explosives and plastics.

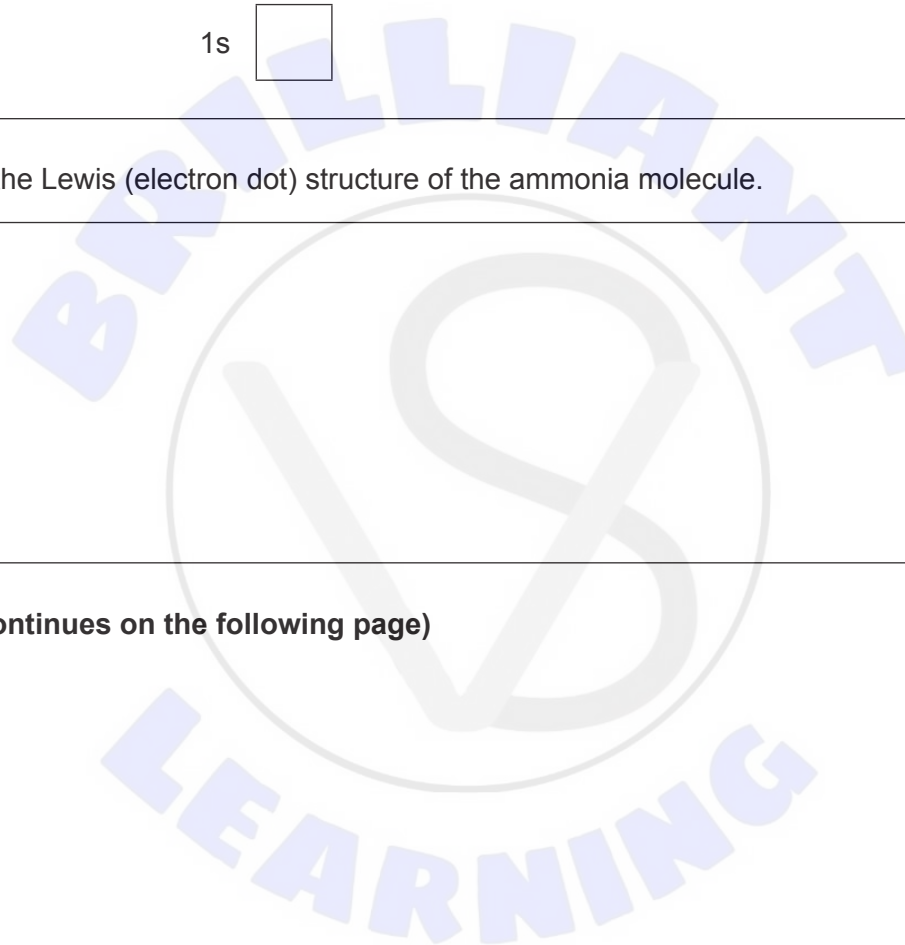
(a) Draw arrows in the boxes to represent the electron configuration of a nitrogen atom. [1]

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2s	<input type="checkbox"/>		
1s	<input type="checkbox"/>		

(b) Draw the Lewis (electron dot) structure of the ammonia molecule. [1]

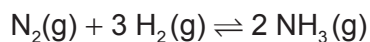


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(Question 2 continued)

(c) Ammonia is produced by the Haber–Bosch process which involves the equilibrium:



(i) Deduce the expression for the equilibrium constant, K_c , for this equation. [1]

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(ii) Explain why an increase in pressure shifts the position of equilibrium towards the products and how this affects the value of the equilibrium constant, K_c . [2]

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(iii) State how the use of a catalyst affects the position of the equilibrium. [1]

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(Question 2 continued)

(d) The effect of temperature on the position of equilibrium depends on the enthalpy change of the reaction.

(i) Determine the enthalpy change, ΔH , for the Haber–Bosch process, in kJ. Use Section 11 of the data booklet. [3]

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(ii) Calculate the enthalpy change, ΔH^\ominus , for the Haber–Bosch process, in kJ, using the following data.

$\Delta H_f^\ominus(\text{NH}_3) = -46.2 \text{ kJ mol}^{-1}$. [1]

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(iii) Suggest why the values obtained in (d)(i) and (d)(ii) differ. [1]

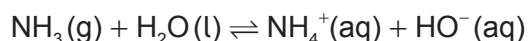
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(Question 2 continued)

(e) Ammonia is soluble in water and forms an alkaline solution:



(i) State the relationship between NH_4^+ and NH_3 in terms of the Brønsted–Lowry theory. [1]

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(ii) Determine the concentration, in mol dm^{-3} , of the solution formed when 900.0 dm^3 of $\text{NH}_3(\text{g})$ at 300.0 K and 100.0 kPa , is dissolved in water to form 2.00 dm^3 of solution. Use sections 1 and 2 of the data booklet. [2]

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(iii) Calculate the concentration of hydroxide ions in an ammonia solution with $\text{pH} = 9.3$. Use sections 1 and 2 of the data booklet. [1]

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(Question 3 continued)

(d) Compound B can also be prepared by reacting an alkene with water.

(i) Draw the structural formula of the alkene required.

[1]

Compound B

$+ \text{H}_2\text{O} \rightarrow$

$$\begin{array}{c} \text{OH} \\ | \\ \text{CH}_3 - \text{C} - \text{CH}_3 \\ | \\ \text{CH}_3 \end{array}$$

(ii) Deduce the structural formula of the repeating unit of the polymer formed from this alkene.

[1]

(e) Deduce what would be observed when Compound B is warmed with acidified aqueous potassium dichromate (VI).

[1]

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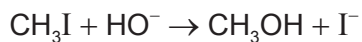
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(Question 3 continued)

(f) Iodomethane is used to prepare CH_3MgI . It can also be converted into methanol:



(i) Identify the type of reaction. [1]

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(ii) Outline the requirements for a collision between reactants to yield products. [2]

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(iii) The polarity of the carbon-halogen bond, C-X, facilitates attack by HO^- . Outline, giving a reason, how the bond polarity changes going down group 17. [1]

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References:

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